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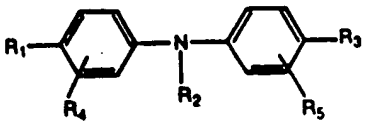
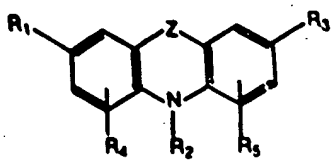
(54) Method for determination of peroxide and test reagent therefor.

(57) A method for the determination of peroxide in a sample is disclosed which comprises reacting the peroxide with a chromogen represented by the general formula (I) or (II)

general formula (I):

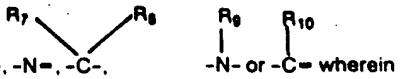


general formula (II):



wherein R_1 and R_3 represent amino, mono- or di-substituted amino, hydroxyl or hydroxyalkyl, R_4 and R_5 represent hydroalkyl, alkylene, acyl, halogen, sulphone, nitro, carboxyl, hydroxyl or hydroxyalkyl, R_2 represents hydrogen,

wherein R_6 represents hydrogen, alkyl, aralkyl, alkylene, aryl or mono- or di-substituted aryl, and -Z- may change to -Z= by resonance and



represents -S-, -O-, -N=, -C-, -N- or -C= wherein R_7, R_8, R_9 and R_{10} have the same significance as R_6 in the presence of heme compound, iodide or bromide and measuring the absorbancy of the reaction solution in the visible ray region. Also disclosed is a test composition for the determination of peroxide which comprises a chromogen as defined above and a compound selected from heme compound, iodide and bromide.

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METHOD FOR DETERMINATION OF PEROXIDE AND TEST REAGENT THEREFOR

The present invention relates to a method and test reagent for the determination of peroxide. More particularly, it relates to a method and test reagent for the determination of peroxide in a sample by reacting the peroxide with a compound which is converted to a pigment by oxidation (hereinafter referred to as "chromogen") in the presence of heme compound, iodide or bromide and measuring the absorbancy of the reaction solution colored in the visible ray region.

The determination of peroxide in vivo is recognized as important for the diagnosis of arteriosclerosis, diabetes mellitus, etc.

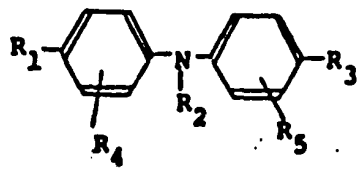
As the methods for the determination of peroxide in a sample, direct methods such as iodide titration method, rohdan iron method, chromatograph method and ultraviolet absorption method and indirect method such as thiobarbituric acid method are known. However, these methods are not satisfactory with respect to sensitivity, and further they require the removal of the substance contained in the sample and affecting the determination.

Recently, a method has been proposed wherein peroxide is determined by reacting peroxide with a chromogen in the presence of a kind of metal compound and by measuring the absorbancy of the reaction solution colored by the formation of pigment. (Japanese Published Unexamined Patent Application Nos. 92391/79 and 23401/80). A simple method which is excellent in sensitivity is in demand.

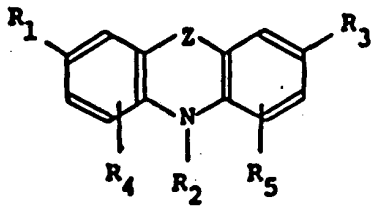
To this end, studies have been made, and it has been found that peroxide and cumene hydroperoxide are determined by reaction with a chromogen represented

by the general formula (I) or (II) below to form a pigment, followed by measurement of the absorbancy of the colored reaction solution in the visible ray region.

Formula (I):



Formula (II):



In the formulae, R_1 and R_3 represent amino, mono- or di-substituted amino, hydroxyl or hydroxyalkyl, R_4 and R_5 represent hydrogen, alkyl, alkylene, acyl, halogen, sulphone, nitro, carboxyl, hydroxyl or hydroxyalkyl, R_2 represents hydrogen, $-\overset{\overset{O}{\|}}{C}-R_6$, $-\overset{\overset{S}{\|}}{C}=R_6$, $\overset{\overset{O}{\|}}{C}-NHR_6$ or $-\overset{\overset{S}{\|}}{C}-NHR_6$ wherein R_6 represents hydrogen, alkyl, aralkyl, alkylene, aryl or mono- or di-substituted aryl, and -Z- may change to -Z= by resonance and represents -S-, -O-, -N=, $-\overset{R_7}{\underset{}{C}}\overset{R_8}{}-$, $-\overset{R_9}{N}-$ or $-\overset{R_{10}}{C}=$ wherein R_7, R_8, R_9 and R_{10} have the same significance as R_6. When Z has three bonds, the position of the double bond of the compound represented by the general formula (II) may change.

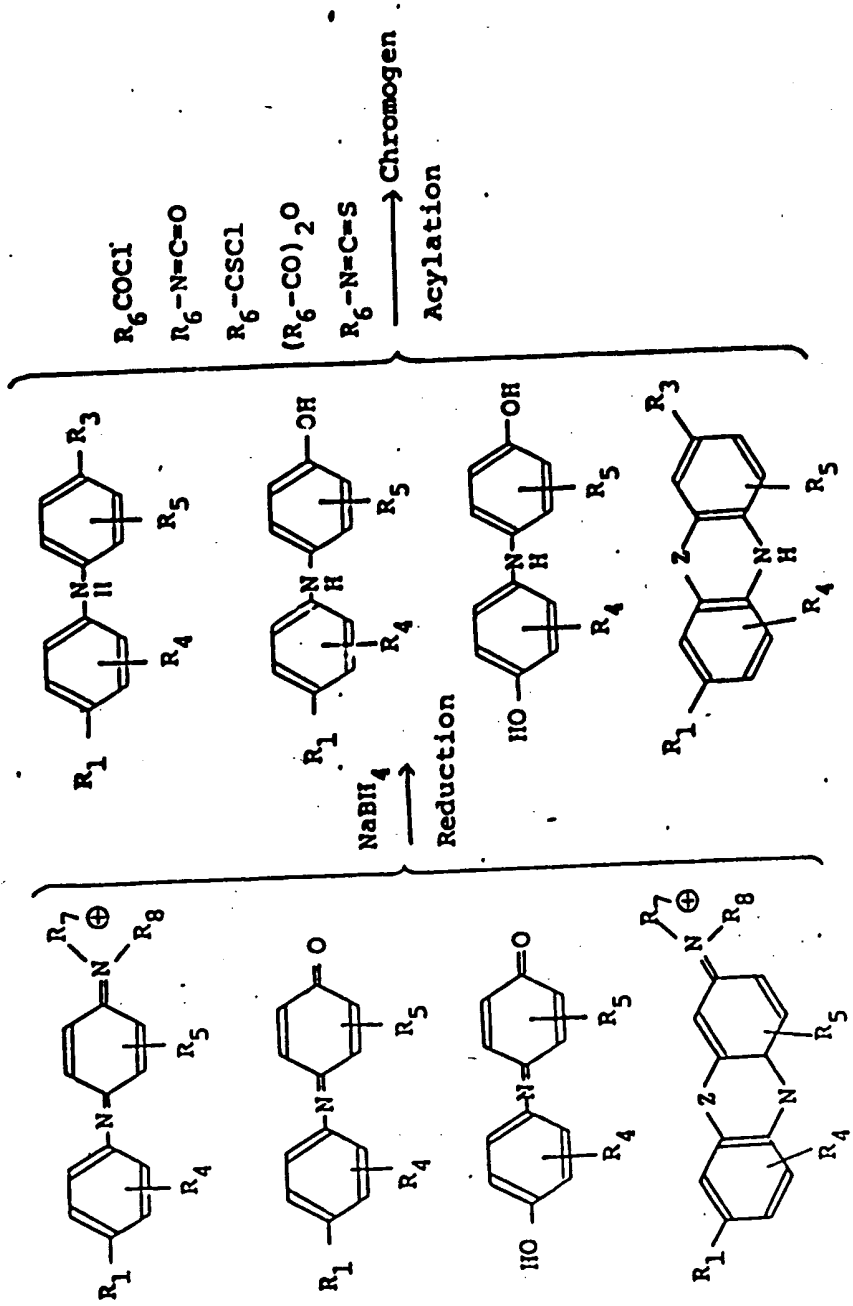
Substituents of substituted amino in R_1 and R_3 include alkyl, alkylene, hydroxyalkyl, acylaminoalkyl

and acyl. As the substituted aryl in R_6, substituted phenyl is exemplified and the substituents include halogen, alkyl, amino, acylamino, and alkoxycarbonylamino. Aryl includes phenyl.

As an aralkyl, phenylalkyl such as benzyl and substituted phenylalkyl such as substituted benzyl are exemplified. The substituents have the same significance as those in the substituted aryl mentioned above.

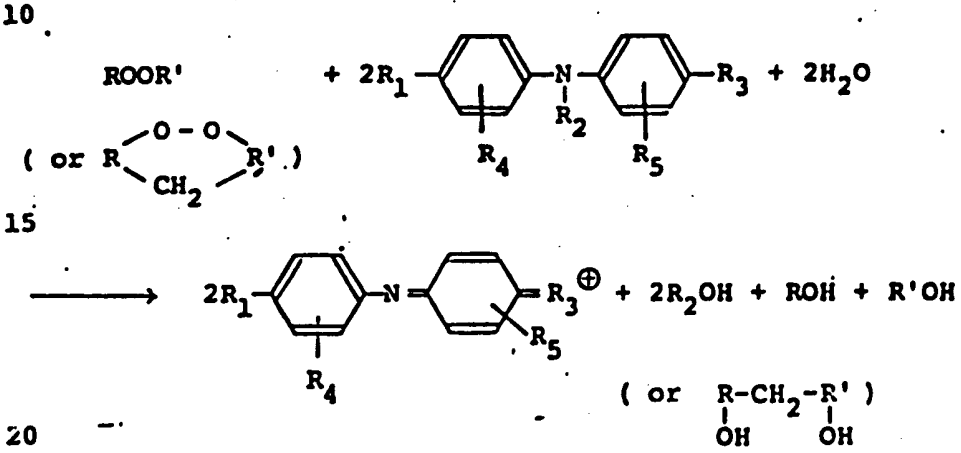
In the above definition, alkyl includes an alkyl group having 1 - 6 carbon atoms and methyl, ethyl, propyl, butyl, pentyl, hexyl and cyclohexyl are exemplified. Acyl includes an acyl group having 2 - 5 carbon atoms, and acetyl and propionyl are exemplified. Alkoxy include an alkoxy group having 1 - 5 carbon atoms and methoxy, ethoxy, propoxy and butoxy are exemplified.

These compounds are generally known and are easily prepared by the methods illustrated by the following reaction formulae.

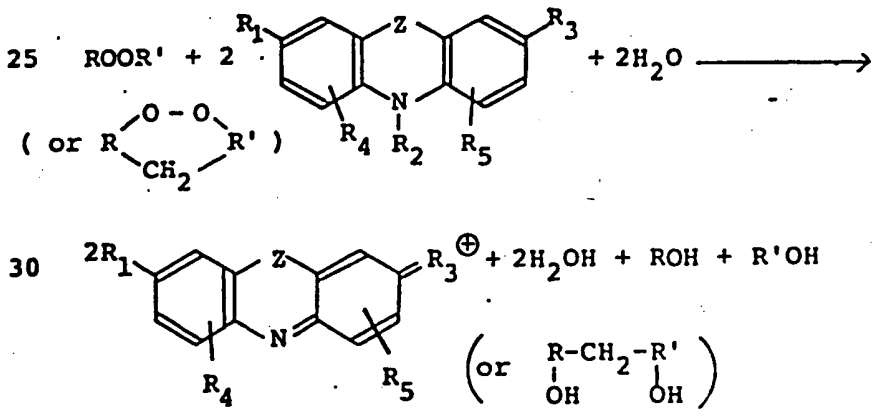

R_6COCl
$R_6-N=C=O$
R_6-CSCl
$(R_6-CO)_2O$
$R_6-N=C=S$
Acylation
Chromogen
$NaBH_4$
Reduction

The principal of the present invention is n the basis of the fact that the reacti n f per xide or cumene hydroperoxide with a chromogen in the presence of heme compound, iodide or bromide proceeds stoichiometrically to form a pigment and the amount of formed pigment is proportional to the amount of peroxide or cumene hydroperoxide in the sample.

The principle is illustrated as follows.

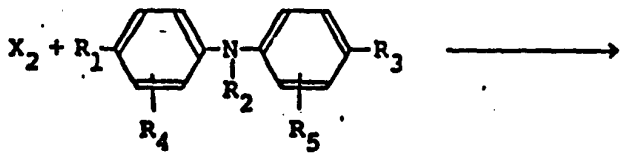


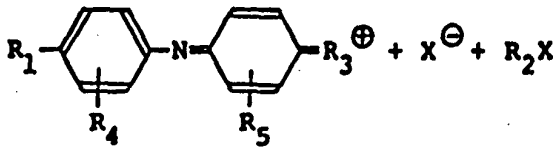
[General formula (I')]



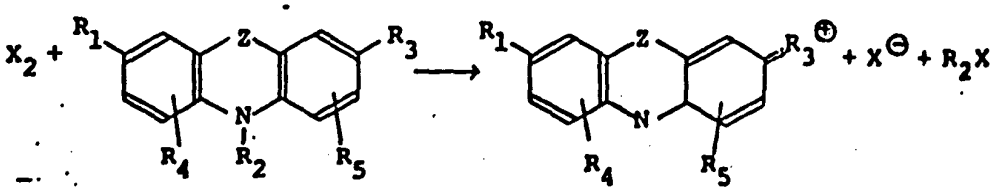
[General formula (II')]

$$ROOR' + 2KX + 2H_2O \longrightarrow ROH + R'OH + 2KOH + X_2$$



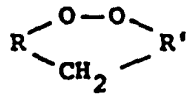


[General formula (I')]



[General formula (II')]

In the above formulae, R_1, R_2, R_3, R_4, R_5 and Z have the same significance as defined above and X is I or Br.

ROOR' and  represent a peroxide.

As is apparent from the above equations, one molecule of the compound represented by the general formula (I') or (II') per one -O-O- group is produced by the reaction and therefore the number of the -O-O- groups in a sample is determined according to the present invention.

The compounds represented by the general formula (I') or (II') are generally known pigments which exhibit a characteristic absorption at the wavelength between 500 - 800 nm and have a large value of

molecular extinction coefficient.

According to the present invention, the present method is applied to the determination of peroxide in a sample such as serum, blood, etc.

In carrying out the present method, a sample is used as itself or after dilution with water, propanol, etc. and if necessary the solution is subjected to centrifugation to remove the substance which may interfere with the measurement and the supernatant is used as a test sample. Usually the sample is used in a concentration of 1 - 500 nmol/ml, preferably 30 - 200 nmol/ml as -O-O- group.

The sample is added to the appropriate buffer solution, preferably, buffer having a pH of 2 - 10. Then to the solution are added, (1) heme compound, iodide or bromide (2) chromogen represented by the general formula (I) or (II) and if necessary, (3) surfactant for promoting the dissolution of peroxide, a chelating reagent such as EDTA for chelating the metal in the sample and sodium chloride for inhibiting the ceruloplasmin activity.

The reaction is carried out at a temperature of 10 - 15°C, preferably 30 - 40°C and usually completes in 5 - 30 minutes. After completion of the reaction, the absorbancy of the reaction solution (E_S) is measured at the characteristic absorption wavelength of the pigment formed from chromogen.

The diluent used for the dilution of sample and the standard compound such as cumene hydroperoxide are subjected to the same procedures as described above to obtain blank absorbancy (E_B) and standard absorbancy (E_{STD}).

The concentration of peroxide (L_p) is calculated by the following equation.

$$L_p = \frac{E_S - E_B}{E_{STD} - E_B} \times A$$

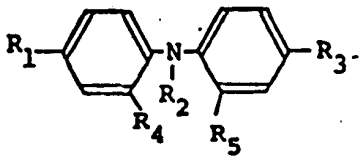
A : the concentration of peroxide in standard solution

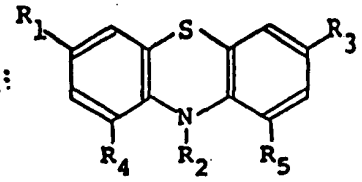
As the heme compound used in the present invention, hemoglobin, myoglobin and iron chlorophyllin are exemplified.

Iodide and bromide include alkaline metal salts such as potassium salt, sodium salt and lithium salt and alkali earth metal salts such as calcium salt, aluminum salt and barium salt of iodine or bromine. The heme compound is used in a concentration of 0.1mg/l-20g/l. Iodide and bromide are usually used in a concentration of 1 - 100 mg/ml. Surfactant, chelating agent and sodium chloride are used in a concentration of 0.001 - 10%. Chromogen is used in a concentration of 0.001 - 1 mg/ml.

As buffers, phosphate buffer, tris-HCl buffer, succinate buffer, citrate buffer, acetate buffer, etc. may be used in a concentration of 0.005-2 mol/l.

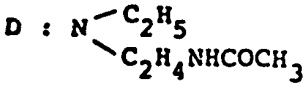
Examples of the chromogen used in the present invention are shown in Table 1. The symbols in Table 1 have the following meaning.

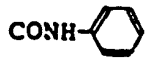
I : 

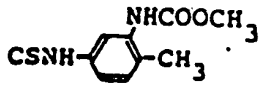
II : 

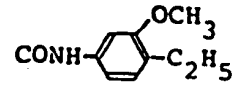
A : $N(CH_3)_2$

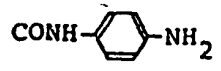
B : $N(C_2H_5)_2$

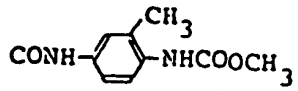
D : 

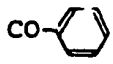
E : 

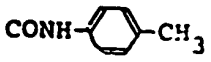
F : 

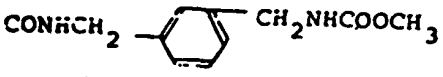
G : 

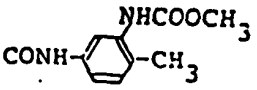
K : 

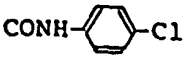
L : 

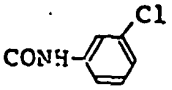
M : 

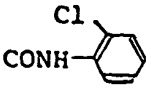
P : 

T : 

Q : 

Y_p : 

Y_m : 

Y_o : 

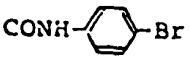
Y_B : 

Table 1

Compound number	Formula	R_1	R_2	R_3	R_4	R_5
1	I	A	H	A	H	H
2	I	NH_2	H	A	H	H
3	I	A	H	A	CH_3	CH_3
4	I	A	E	B	H	H
5	I	D	F	D	H	H
6	I	D	H	D	CH_3	CH_3
7	I	A	G	OH	CH_3	H
8	I	OH	K	OH	CH_3	H
9	I	A	L	B	H	OH
10	II	A	H	A	H	H
11	II	A	M	A	H	H
12	II	B	P	D	CH_3	H
13	II	D	Q	D	H	H
14	II	D	Q	OH	CH_3	CH
15	II	A	Y_P	A	H	H
16	II	A	Y_O	A	H	H
17	II	A	Y_m	A	H	H
18	II	A	Y_B	A	H	H
19	II	A	E	A	H	H
20	II	A	T	A	H	H

The degree of color development, the stability of color and the influence of the components in serum on the determination value when Compound (I) is used as a chromogen, are illustrated by the following experiment.

Experiment 1

Compound Nos. 1-20 in the amount indicated in Table 2 are dissolved in 1 ml of dimethylformamide (hereinafter referred tc as "DMF"). The test reagent is prepared by adding 0.1 g Triton X-100, DMF solution of Compound Nos. 1-20, 1 g of EDTA and 6.7 mg of hemoglobin t 100 ml of 0.1 M phosphate buffer (pH 5.0).

3 ml of the reagent s luti n is poured into a test tube for each compound and linolic acid (A) is added thereto. The reaction is carried out at 37°C for 30 minutes and the absorbancy of the reaction solution (E_S) is measured. The blank absorbancy (E_B) is measured by repeating the above procedures without the addition of linolic acid.

As a control, the absorbancies (E_{SC} and E_{BC}) are measured using 4-amino antipyrine (hereinafter referred to as 4 AA) and m-methyl-(N-ethyl, N-acetoaminoethyl) aniline (hereinafter referred to as EMAE) as coloring reagent, and the degree of color development of test compound is calculated from the following equation defining the degree of color development of control as 100.

$$\text{Degree of color development} = \frac{E_S - E_B}{E_{SC} - E_{BC}} \times 100$$

The stability of color is determined as follows. The reaction solution is further incubated at 37°C for 30 minutes and $E_S - E_B$ is calculated. "S_0" means that the value of $E_S - E_B$ was not changed by this incubation and "S_1" means that the change of the value is 10% or less. The influence of bilirubin and vitamin C in the sample on the determination value is indicated by the value M. The value M is determined by repeating the above experiment using the sample containing 4 µg/3 ml of bilirubin or 2 µg/3 ml of vitamin C, measuring the absorbancy ($E_{S'}$) and calculating from the following equation:

$$M = \frac{E_S - E_{S'}}{E_S - E_B} \times 100 \ (\%)$$

The symbol (-) means that the value (M) is 3% or less. The symbol (±) means that the value is 3-6% and (+) for 6-20% and (++) for 20% or more.

Table 2

Chromogen		Degree of color development	Inhibitor		Stability
No.	Amount mg		B_i	V_C	
1	10.0	538	-	-	S_0
2	10.0	313	-	-	S_1
3	12.0	277	-	±	S_0
4	14.7	495	-	+	S_0
5	23.4	532	-	±	S_0
6	16.8	575	-	-	S_0
7	16.4	251	+	±	S_1
8	14.3	304	+	+	S_1
9	19.2	527	-	±	S_0
10	11.2	572	-	±	S_0
11	15.0	470	-	±	S_0
12	21.4	581	-	±	S_0
13	24.8	519	-	±	S_0
14	22.1	294	±	+	S_0
15	3.1	506	-	±	S_0
16	3.1	311	±	±	S_0
17	3.1	547	-	±	S_0
18	10.0	523	-	±	S_0
19	3.3	536	-	±	S_0
20	3.3	478	-	±	S_0
Control	4 AA : 6.7 EMAE : 26.6	100	-	±	S_0
Comparison 1	4 AA : 6.7 Phenol: 33	43.4	++	+	S_0
Comparison 2	M-T method	210	++	+	S_0

B_i : bilirubin

V_C : vitamin C

M-T : malon diald hyde-thiobarbituric acid

For comparison, 4 AA-phenol or malon dialdehydethiobarbituric acid is used as a chromogen and th results are shown in Table 2.

Another aspect of the present invention is to provide a test composition for the determination of peroxide which comprises a chromogen represented by the general formula (I) or (II), a compound selected from heme compound, iodide and bromide and a buffer. The composition may contain a surfactant, chelating reagent and sodium chloride.

The composition may be used in various forms. For example, the ingredients may be mixed in liquid form or powder form.

Certain specific embodiments of the present invention are illustrated by the following representative examples.

Example 1

The test reagent is prepared by adding 0.1 g of Triton X-100, 1 ml of DMF solution containing 10 mg of Compound 1, 5.6 mg of hemoglobin and 1 g of EDTA to 100 ml of 0.1 M phosphate buffer (pH 5.0).

As the sample containing peroxide, 1 ml of linolic acid (A) and 1 ml of linolenic acid (B) are respectively diluted with isopropanol to make 100 ml of a solution.

20 µl of the test sample is added to 3 ml of the test reagent and incubated at 37°C with stirring. The absorbancy of the reaction solution at 728 nm is monitored for 10 minutes from the start of the reaction. The absorbancy reaches equilibrium within about 5 minutes.

The same procedures as described above are repeated except that cumene hydroperoxide is used as a standard compound and the absorbancy of the reaction solution at 728 nm is measured about 10 minutes from the start of the reaction. The standard curve between the

absorbancy and th concentrati n of peroxide is prepared by repeating the above procedure varying the concentration of cumene hydroperoxide.

The peroxide values for the samples containing linolic acid (A) and linolenic acid (B) are calculated from the standard curve to obtain 37 for (A) and 29.5 for (B).

For comparison, peroxide values of the samples are determined according to the known iodide titration method to obtain 35.1 for (A) and 21.3 for (B).

The present method and the known method described above are repeated five times for samples (A) and (B). The coefficient of variation by the present method is 0.1% for (A) and 0.15% for (B) and that by the known method is 10.5% for (A) and 12.3% for (B).

Example 2

The samples indicated in Table 3 are dissolved in water or isopropanol in a ratio of 10% (V/V). 50 µl of each solution is added to 3 ml of the test reagent of Example 1. The reaction is carried out under the same conditions as in Example 1 and the absorbancy of the reaction solution is measured.

The results are shown in Table 3.

Table 3

Sample	Solvent	The amount of peroxide *1 (nmol/g)	C.V (%) *2
1	Distilled water	390.7	0.31
2	"	274.4	0.52
3	"	135.9	0.86
4	"	183.1	0.42
5	Isopropanol	15.2	3.21

1 : Emulgen 404 (a non-ionic surfactant, product of Kao Atlas Co., Ltd.)

2 : Emul 20T (anionic surfactant, product of Kao Atlas Co., Ltd.)

3 : Quartamin 86P (cationic surfactant, product of Kao Atlas Co., Ltd.)

4 : Tetrahydrofuran

5 : Ethylether

*1 : The value is the average of five measurements.

*2 : CV : coefficient of variation

Example 3

In this example, 0.2 ml of normal and patient serum are added to 4 ml of isopropanol and the solutions are subjected to centrifugation at 2,000 r.p.m. for 5 minutes. To 0.5 ml of the supernatant is added 3 ml of the test reagent of Example 1, and the mixture is incubated at 37°C for 10 minutes. The absorbancy of the reaction solution (E_S) is measured at 728 nm.

The same procedures as described above are repeated for 0.5 ml of 200 nmol/ml cumene hydroperoxide isopropanol solution and 0.5 ml of isopropanol to obtain the absorbancies E_{STD} and E_B. The amount of peroxide (L_p) in serum calculated from the following equation is 15.2 nmol/ml for normal and 96.3 nmol/ml for patient.

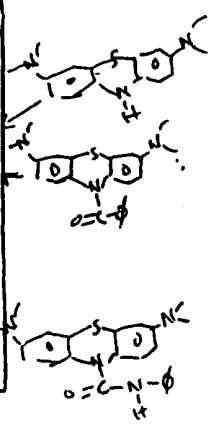
$$L_p(\text{nmol/ml}) = \frac{E_S - E_B}{E_{STD} - E_B} \times 200$$

Example 4

The same procedures as described in Example 1 are repeated except that the compounds indicated in Table 4 instead of Compound 1 are used as chromogen and the peroxide values (PA) for sample (A) and (PB) for sample (B) are determined. The results are shown in Table 4.

Table 4

Chromogen	Wavelength (nm)	Reaction time (min.)	PA (nmol/Kg)	PB (nmol/Kg)
2	700	< 5	38.5	30.4
3	720	< 5	38.3	29.9
4	720	15	36.9	25.1
5	730	15	37.2	26.2
6	730	< 5	38.9	28.9
7	600	20	40.3	30.5
8	600	20	36.1	26.4
9	720	15	39.2	29.8
10	665	< 5	38.5	28.7
11	665	20	37.1	26.8
12	670	15	37.2	27.0
13	670	16	36.8	26.1
14	620	20	35.9	26.5
19	665	< 5	39.0	27.3



Example 5

The same procedures as described in Example 1 are repeated except that 0.56 mg of sodium salt of iron chlorophyllin or 5.6 mg of myoglobin (Sigma Co.) is used instead of hemoglobin. The peroxide value obtained using iron chlorophyllin is 38.2 for sample (A) and 29.3 for sample (B) and that obtained using myoglobin is 37.8 for sample (A) and 28.5 for sample (B).

Example 6

In this example, 0.1 g of Triton X-100, 1 ml of DMF solution containing 10 mg of Compound 1, 1 g of potassium iodide and 1 g of EDTA are dissolved in 100 ml of 0.1 M phosphate buffer (pH 4.0) and the solution is used as test reagent.

The same procedures as described in Example 1 are repeated using the test reagent for 50 µl of sample (A) or (B). The peroxide value is 37.5 for sample (A) and 28.7 for sample (B).

Example 7

In this example, as the standard peroxide, linolic acid is oxidized with air at 23°C for 72 hours according to the method described in Canad. J. Biochem. 47, 485 (1969). The oxidation product is subjected to extraction with a solvent system of petroleum ether/67% methanol - 33% water. The layer of methanol - water is concentrated to obtain an oily matter. The oily matter is subjected to thin layer chromatography using silica gel and hexane - ether - acetic acid (60 : 40 : 1) as a developer. Silica gel showing an Rf value of 0.23 is subjected to elution using ethanol and the eluate is concentrated to obtain a racemic mixture of equimolar amounts of linolic acid having -OOH at the 9-position and linolic acid having -OOH at the 13-position. 310.4 mg of the obtained mixture is dissolved in 1 ℓ of isopropanol to obtain 1 µmol/ml solution (hereinafter referred to as test solution).

The test reagent is prepared by adding 0.1 g of Triton X-100, 1 ml of DMF solution containing 10 mg of Compound 20, 5.6 mg of hemoglobin and 1 g of EDTA to 100 ml of 0.1 M phosphate buffer (pH 5.0).

The test solution is diluted ten-fold with isopropanol. 100 µl of the diluted test solution is added to 3 ml of the test reagent in a test tube and 100 µl of isopropanol is added to 3 ml of the test reagent in another test tube. The mixtures are incubated at 37°C for ten minutes and the absorbancies of the reaction solutions are measured at 666 µm to obtain O.D. values of 0.308 and 0.085 respectively. The increase in absorbancy by the addition of oxidized linolic acid is calculated as 0.223.

Then, 391.9 mg of Methylene Blue is dissolved in 1 ℓ of water and the solution is diluted ten-fold with water. 100 µl of the blue colored solution is added to 0.1 M phosphate buffer (pH 5.0) and the absorbancy of the solution is measured to obtain a value of 0.225.

Then, 100 µl of the test solution having the concentration indicated in Table 5 is added to 3 ml of the test reagent and the mixture is incubated at 37°C for ten minutes. The absorbancy (E_N) of the reaction solution is measured at 666 nm. As a blank, the same procedures as described above are repeated except using isopropanol instead of the test solution and the absorbancy (E_B) is measured at 666 nm. The results are shown in Table 5. As is apparent from the table, the concentration of oxidized linolic acid is proportional to the value of $E_N - E_B$.

Table 5

Concentration of linolic acid (nmol/ml)	0	25	50	100	150
$E_N - E_B$	0	0.057	0.113	0.227	0.340

Example 8

In this example, 100 µl of isopropanol solution of cumene hydroperoxide having the concentration indicated in Table 6 is added to 3 ml of 50 mM citrate buffer (pH 4.7) containing 1 mg/ml Emulgen 106 (non-ionic surfactant, product of Kao Atras Co., Ltd.), 5 mg/dl hemoglobin and 2.5 mg/dl Compound 18.

The mixture is incubated at 37°C for 30 minutes. The absorbancy of the reaction solution (E_N) is measured at 666 nm.

As a blank test, the same procedures as described above are repeated except using isopropanol instead of cumene hydroperoxide solution and the absorbancy (E_B) is measured.

The results are shown in Table 6. As is apparent from the table, the concentration of cumene hydroperoxide is proportional to the value of $E_N - E_B$.

Table 6

Concentration of cumene hydroperoxide (nmol/ml)	0	25	50	100	150
$E_N - E_B$	0	0.056	0.113	0.225	0.336

Reference Example 1

In this example, 1 g of Methylene Blue is dissolved in 100 ml of water and 1 g of sodium borohydride is added little by little to proceed the reduction. When the precipitate of leuco base is deposited and the solution is discolored, 20 ml of chloroform is added and the mixture is vigorously stirred to extract leuco base.

The chloroform layer is filtered through filter paper, dehydrated and desalted. Then, 2 ml of phenyl isocyanate is added and the mixture is subjected to reaction at room temperature for 24 hours.

After completion of the reaction, methanol is added to remove excess isocyanate and the mixture is stirred at room temperature for 3 hours.

The mixture is subjected to column chromatography using silica gel having the size of 60 - 80 mesh (product of Kanto Kagaku Co., Ltd.) and using chloroform as a developer to obtain Compound 19 having a melting point of 100 - 115°C.

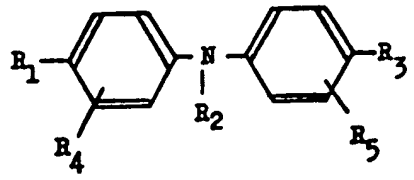
Reference Example 2

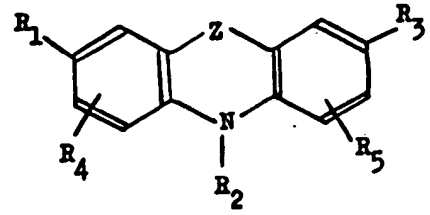
The same procedures as described in Reference Example 1 are repeated except that o-, m- or p- chlorophenyl isocyanate or p-bromophenyl isocyanate is used instead of phenylisocyanate to obtain Compound 16 (oil

form), Compound 17 (m.p. 73 - 77°C), Compound 15 (m.p. 76 - 83°C) and Compound 18 (m.p. 80 - 90°C), respectively.

CLAIMS

1. A method for the determination of peroxide in a sample which comprises reacting the peroxide with a chromogen in the presence f a heme compound, a bromide or an iodide, and measuring the absorbancy of the reaction solution in the visible ray region, characterised in that there is used as said chromogen a compound of the general formula (I) or (II)

I 

II 

wherein R_1 and R_3 represent amino, mono- or di- substituted amino, hydroxyl or hydroxyalkyl, R_4 and R_5 represent hydrogen, alkyl, alkylene, acyl, halogen, sulphone, nitro, carboxyl, hydroxyl or hydroxyalkyl, R_2 represents hydrogen, $-\overset{\overset{O}{\|}}{C}-R_6$, $-\overset{\overset{S}{\|}}{C}=R_6$, $-\overset{\overset{O}{\|}}{C}-NHR_6$ or $-\overset{\overset{S}{\|}}{C}-NHR_6$

wherein R_6 represents hydrogen, alkyl, aralkyl, alkylene, aryl or mono- or di-substituted aryl, and -Z- may change to -Z= by resonance and represents -S-, -O-, -N=, $-\overset{R_7}{\underset{}{C}}\overset{R_8}{}-$, $-\overset{R_9}{N}-$ or $-\overset{R_{10}}{C}=$ wherein R_7, R_8, R_9 and R_{10} have the same significance as R_6.

2. A method according to claim 1, characterised in that the chromogen is used in combination with a heme compound selected from hemoglobin, myoglobin and iron chlorophyllin.

3. A method according to claim 1, characterised in that the chromogen is used in combination with an iodide or bromide selected from alkali metal and alkaline earth metal iodides or bromides.

4. A test composition for the determination of peroxide in a sample, characterised in that it comprises a chromogen as defined in claim 1 in admixture with a heme compound, an iodide or a bromide.

5. A composition according to claim 4 comprising said chromogen in admixture with a heme selected from hemoglobin, myoglobin and iron chlorophyllin.

6. A composition according to claim 4 comprising said chromogen in admixture with an alkali metal or alkaline earth metal bromide or iodide.

European Patent Office

EUROPEAN SEARCH REPORT

Application number EP 81 30 1626

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl.)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
X	US - A - 3 654 179 (R. BAUER) * Column 1, lines 50-65; claims 1 and 3 * ---	1,2,4, 5	G 01 N 33/52 C 12 Q 1/28
	GB - A - 2 002 517 (MILES LABORATORIES INC.) * Page 1, lines 33-42; claims 1-8 * ---	1	
	GB - A - 1 400 897 (CALBIOCHEM AG) * Page 1, lines 20-25; claims 3, 4,7 * ---	1	TECHNICAL FIELDS SEARCHED (Int. Cl.)
	CHEMICAL ABSTRACTS, vol. 80, no. 12, March 25, 1974, page 482, abstract 66505b, COLUMBUS, OHIO (US) H. PUZANOWSKA-TARASIEWICZ: "Phenothiazinederivatives as new indicators in chemical analysis. III. Bromatometric determination of hydroquinone and ascorbic acid", & Chem. Anal. (Warsaw) vol. 18, no. 4, 1973, pages 753-762. * The whole abstract * ---	1,3	G 01 N 33/52 33/66 33/72 C 12 Q 1/26 1/28 1/54 1/62
	CHEMICAL ABSTRACTS, vol. 89, no. 14, October 2, 1978, page 888, abstract 122422x, COLUMBUS, OHIO (US) N.V. RAO et al.: "Oxazine dyes as new iodometric indicators", & J. Indian Chem. Soc., vol. 55, no. 2, 1978, pages 200-203 * The whole abstract * ---	1,3	CATEGORY OF CITED DOCUMENTS X: particularly relevant A: technological background O: non-written disclosure P: intermediate document T: theory or principle underlying the invention E: conflicting application D: document cited in the application L: citation for other reasons
	./.		&: member of the same patent family, corresponding document

X The present search report has been drawn up for all claims

Place of search	Date of completion of the search	Examiner
The Hague	30.06.1981	GRIFFITH

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EUROPEAN SEARCH REPORT

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DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl.³)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
A	US - A - 3 791 988 (DIETER JOSEF)		
A	GB - A - 1 473 945 (GENERAL ELECTRIC Co.)		
A	DE - A - 2 110 342 (KERNFORSCHUNGS-ANLAGE JULICH GmbH)		
	---------		TECHNICAL FIELDS SEARCHED (Int. Cl.³)